

SELAQUI INTERNATIONAL SCHOOL
HOLIDAY HOMEWORK
SUMMER VACATION 2019-20
(CLASS X)

ENGLISH (XA)

1. Visit to nearest post office and complete the assignment as discussed in the classroom.
2. Watch Invictus and complete assignment as discussed in the classroom.

ENGLISH (XB)

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Watch the movie 'The Story of My Life' on the link mentioned below

<https://www.youtube.com/watch?v=-3kqkHT3HzM>

Write Character Sketches for the following characters –

1. Helen
2. Anne Sullivan
3. Arthur H. Keller
4. Alexander Graham Bell

HINDI

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1. विद्यालय में आयोजित की जाने वाली किन्हीं 5 गतिविधियों पर सूचना लेखन कीजिए ।
2. समाचार पत्र-पत्रिकाओं से कोई 10 विज्ञापन (सभी अलग तरह के) काट कर गृहकार्य पुस्तिका में चिपकाएँ व उनके आधार उनसे मिलते-जुलते 5 विज्ञापन तैयार करें ।
विज्ञापन :
पर्यावरण अथवा सामाजिक समस्या पर विज्ञापन , कोई नया/पुराना सामान अथवा उत्पाद (product) बेचने सम्बन्धी विज्ञापन, जागरूक (aware) करने संबंधी विज्ञापन, आवश्यकता संबंधी विज्ञापन आदि ।
3. किन्हीं तीन विषयों पर संवाद लेखन कीजिए ।
4. निम्नलिखित विषयों के आधार पर 100 से 120 शब्दों में अनुच्छेद लेखन कीजिए
क) पर्यावरण संबंधी ख) तकनीक और विज्ञान ग) कोई सामाजिक समस्या घ) कृषि,
किसान और गाँव ड) अनुशासन,समय आदि ।

FRENCH

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Faites une brochure touristique de La Corse avec les images et toutes les caractéristiques d'une brochure informative et publicitaire en combinant caractéristiques géographiques, les infrastructures disponibles et les différentes activités pour les visiteurs. (au format A 6)

Imaginez un dialogue avec votre ami francophone sur le climat en Inde pendant les mois d'été. (60 mots)

SCIENCE

PHYSICS

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Questions: 1 to 18 are multiple choice questions and 19 to 35 are short and long answer type questions.

The Physics worksheet can also be downloaded from the link given below.

NCERT EXEMPLAR link for downloading the worksheet <http://ncert.nic.in/ncerts/l/jeeep112.pdf>

CHAPTER 12
Electricity

Multiple Choice Questions

1. A cell, a resistor, a key and an ammeter are arranged as shown in the circuit diagrams of Figure 12.1. The current recorded in the ammeter will be

Fig. 12.1

(a) maximum in (i)
(b) maximum in (ii)
(c) maximum in (iii)
(d) the same in all the cases

2. In the following circuits (Figure 12.2), heat produced in the resistor or combination of resistors connected to a 12 V battery will be

Fig. 12.2

(a) same in all the cases
(b) maximum in case (i)
(c) maximum in case (ii)
(d) maximum in case (iii)

3. Electrical resistivity of a given metallic wire depends upon
- its length
 - its thickness
 - its shape
 - nature of the material
4. A current of 1 A is drawn by a filament of an electric bulb. Number of electrons passing through a cross section of the filament in 16 seconds would be roughly
- 10^{20}
 - 10^{16}
 - 10^{18}
 - 10^{23}
5. Identify the circuit (Figure 12.3) in which the electrical components have been properly connected.

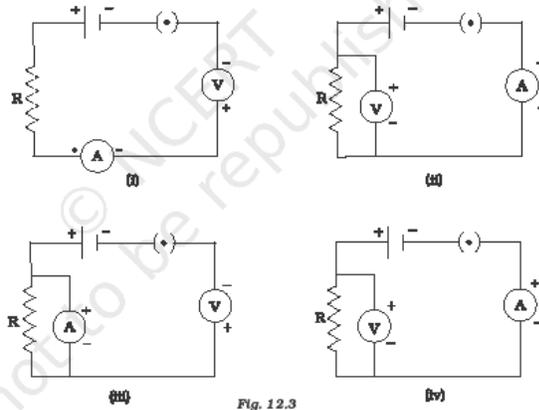
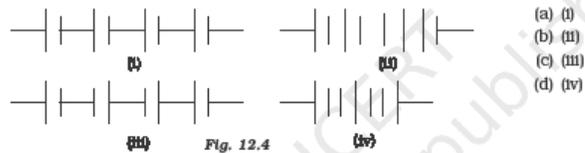


Fig. 12.3

- (i)
- (ii)
- (iii)
- (iv)

6. What is the maximum resistance which can be made using five resistors each of $1/5 \Omega$?
- (a) $1/5 \Omega$
 (b) 10Ω
 (c) 5Ω
 (d) 1Ω
7. What is the minimum resistance which can be made using five resistors each of $1/5 \Omega$?
- (a) $1/5 \Omega$
 (b) $1/25 \Omega$
 (c) $1/10 \Omega$
 (d) 25Ω
8. The proper representation of series combination of cells (Figure 12.4) obtaining maximum potential is



9. Which of the following represents voltage?

- (a) $\frac{\text{Work done}}{\text{Current} \times \text{Time}}$
 (b) $\text{Work done} \times \text{Charge}$
 (c) $\frac{\text{Work done} \times \text{Time}}{\text{Current}}$
 (d) $\text{Work done} \times \text{Charge} \times \text{Time}$

10. A cylindrical conductor of length l and uniform area of cross-section A has resistance R . Another conductor of length $2l$ and resistance R of the same material has area of cross section
- (a) $A/2$
 (b) $3A/2$
 (c) $2A$
 (d) $3A$

11. A student carries out an experiment and plots the V-I graph of three samples of nichrome wire with resistances R_1 , R_2 and R_3 respectively (Figure.12.5). Which of the following is true?

- (a) $R_1 = R_2 = R_3$
(b) $R_1 > R_2 > R_3$
(c) $R_3 > R_2 > R_1$
(d) $R_2 > R_3 > R_1$

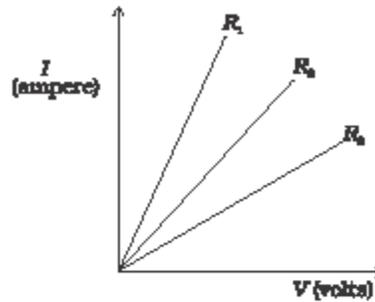


Fig. 12.5

12. If the current I through a resistor is increased by 100% (assume that temperature remains unchanged), the increase in power dissipated will be
- (a) 100 % (b) 200 %
(c) 300 % (d) 400 %
13. The resistivity does not change if
- (a) the material is changed
(b) the temperature is changed
(c) the shape of the resistor is changed
(d) both material and temperature are changed
14. In an electrical circuit three incandescent bulbs A, B and C of rating 40 W, 60 W and 100 W respectively are connected in parallel to an electric source. Which of the following is likely to happen regarding their brightness?
- (a) Brightness of all the bulbs will be the same
(b) Brightness of bulb A will be the maximum
(c) Brightness of bulb B will be more than that of A
(d) Brightness of bulb C will be less than that of B
15. In an electrical circuit two resistors of 2Ω and 4Ω respectively are connected in series to a 6 V battery. The heat dissipated by the 4Ω resistor in 5 s will be
- (a) 5 J
(b) 10 J
(c) 20 J
(d) 30 J
16. An electric kettle consumes 1 kW of electric power when operated at 220 V. A fuse wire of what rating must be used for it?
- (a) 1 A
(b) 2 A
(c) 4 A
(d) 5 A

17. Two resistors of resistance $2\ \Omega$ and $4\ \Omega$ when connected to a battery will have
- same current flowing through them when connected in parallel
 - same current flowing through them when connected in series
 - same potential difference across them when connected in series
 - different potential difference across them when connected in parallel
18. Unit of electric power may also be expressed as
- volt ampere
 - kilowatt hour
 - watt second
 - joule second

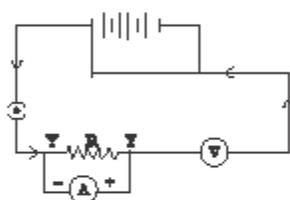


Fig. 12.6

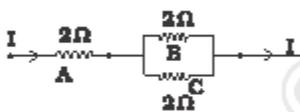


Fig. 12.7

Short Answer Questions

19. A child has drawn the electric circuit to study Ohm's law as shown in Figure 12.6. His teacher told that the circuit diagram needs correction. Study the circuit diagram and redraw it after making all corrections.
20. Three $2\ \Omega$ resistors, A, B and C, are connected as shown in Figure 12.7. Each of them dissipates energy and can withstand a maximum power of 18W without melting. Find the maximum current that can flow through the three resistors?
21. Should the resistance of an ammeter be low or high? Give reason.
22. Draw a circuit diagram of an electric circuit containing a cell, a key, an ammeter, a resistor of $2\ \Omega$ in series with a combination of two resistors ($4\ \Omega$ each) in parallel and a voltmeter across the parallel combination. Will the potential difference across the $2\ \Omega$ resistor be the same as that across the parallel combination of $4\ \Omega$ resistors? Give reason.
23. How does use of a fuse wire protect electrical appliances?
24. What is electrical resistivity? In a series electrical circuit comprising a resistor made up of a metallic wire, the ammeter reads 5A . The reading of the ammeter decreases to half when the length of the wire is doubled. Why?

25. What is the commercial unit of electrical energy? Represent it in terms of joules.
26. A current of 1 ampere flows in a series circuit containing an electric lamp and a conductor of $5\ \Omega$ when connected to a 10 V battery. Calculate the resistance of the electric lamp.
- Now if a resistance of $10\ \Omega$ is connected in parallel with this series combination, what change (if any) in current flowing through $5\ \Omega$ conductor and potential difference across the lamp will take place? Give reason.
27. Why is parallel arrangement used in domestic wiring?
28. B_1 , B_2 and B_3 are three identical bulbs connected as shown in Figure 12.8. When all the three bulbs glow, a current of 3A is recorded by the ammeter A.
- What happens to the glow of the other two bulbs when the bulb B_1 gets fused?
 - What happens to the reading of A_1 , A_2 , A_3 and A when the bulb B_2 gets fused?
 - How much power is dissipated in the circuit when all the three bulbs glow together?

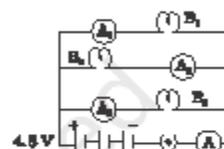


Fig. 12.8

Long Answer Questions

29. Three incandescent bulbs of 100 W each are connected in series in an electric circuit. In another circuit another set of three bulbs of the same wattage are connected in parallel to the same source.
- Will the bulb in the two circuits glow with the same brightness? Justify your answer.
 - Now let one bulb in both the circuits get fused. Will the rest of the bulbs continue to glow in each circuit? Give reason.
30. State Ohm's law? How can it be verified experimentally? Does it hold good under all conditions? Comment.
31. What is electrical resistivity of a material? What is its unit? Describe an experiment to study the factors on which the resistance of conducting wire depends.
32. How will you infer with the help of an experiment that the same current flows through every part of the circuit containing three resistances in series connected to a battery?
33. How will you conclude that the same potential difference (voltage) exists across three resistors connected in a parallel arrangement to a battery?

34. What is Joule's heating effect? How can it be demonstrated experimentally? List its four applications in daily life.
35. Find out the following in the electric circuit given in Figure 12.9
- Effective resistance of two $8\ \Omega$ resistors in the combination
 - Current flowing through $4\ \Omega$ resistor
 - Potential difference across $4\ \Omega$ resistance
 - Power dissipated in $4\ \Omega$ resistor
 - Difference in ammeter readings, if any.

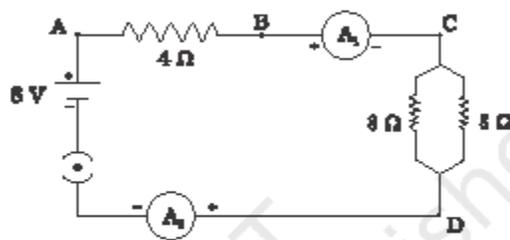


Fig. 12.9

CHEMISTRY

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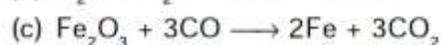
1. Write the balanced chemical equations for the following reactions and identify the type of reaction in each case.

- (a) Thermite reaction, iron (III) oxide reacts with aluminum and gives molten iron and aluminum oxide.
- (b) Magnesium ribbon is burnt in an atmosphere of nitrogen gas to form solid magnesium nitride.
- (c) Chlorine gas is passed in an aqueous potassium iodide solution to form potassium chloride solution and solid iodine.
- (d) Ethanol is burnt in air to form carbon dioxide, water and releases heat.

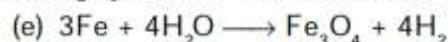
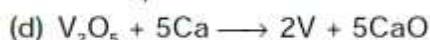
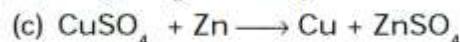
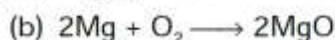
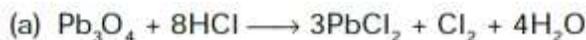
2. Which among the following changes are exothermic or endothermic in nature?

- (a) Decomposition of ferrous sulfate
- (b) Dilution of sulfuric acid
- (c) Dissolution of sodium hydroxide in water
- (d) Dissolution of ammonium chloride in water

3. Identify the reducing agent in the following reactions



4. Identify the oxidizing agent (oxidant) in the following reactions

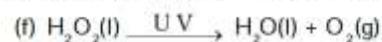
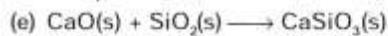
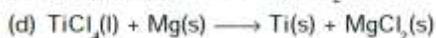
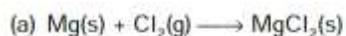


5. Ferrous sulfate decomposes

with the evolution of a gas having a characteristic odor of burning sulfur. Write the chemical reaction involved and identify the type of reaction.

6. Why do we store silver chloride in dark colored bottles?

7. Balance the following chemical equations and identify the type of chemical reaction.



8. A magnesium ribbon is burnt in oxygen to give a white compound X accompanied by emission of light. If the burning ribbon is now placed in an atmosphere of nitrogen, it continues to burn and forms a compound Y.

(a) Write the chemical formulae of X and Y.

(b) Write a balanced chemical equation, when X is dissolved in water.

9. On heating blue colored powder of copper (II) nitrate in a boiling tube, copper oxide (black), oxygen gas and a brown gas X is formed.

(a) Write a balanced chemical equation of the reaction.

(b) Identify the brown gas X evolved.

(c) Identify the type of reaction.

(d) What could be the pH range of aqueous solution of the gas X?

10. Give the characteristic tests for the following gases

(a) CO_2

(b) SO_2

(c) O_2

(d) H_2

11. What happens when a piece of

(a) zinc metal is added to copper sulfate solution?

(b) aluminium metal is added to dilute hydrochloric acid?

(c) silver metal is added to copper sulfate solution?

Also, write the balanced chemical equation if the reaction occurs

12. What happens when zinc granules are treated with dilute solution of H_2SO_4 , HCl , HNO_3 , NaCl and NaOH , also write the chemical equations if reaction occurs.

13. On adding a drop of barium chloride solution to an aqueous solution of sodium sulfite, a white precipitate is obtained.

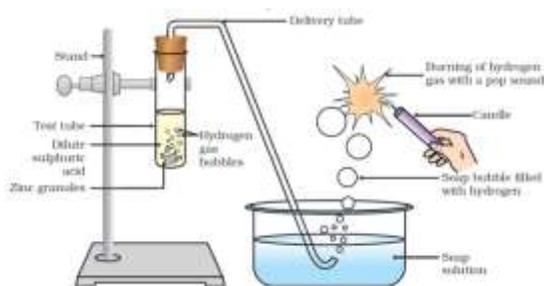
(a) Write a balanced chemical equation of the reaction involved

(b) What other name can be given to this precipitation reaction?

(c) On adding dilute hydrochloric acid to the reaction mixture, white precipitate disappears. Why?

14. You are provided with two containers made up of copper and aluminium. You are also provided with solutions of dilute HCl , dilute HNO_3 , ZnCl_2 and H_2O . In which of the above containers these solutions can be kept?

15. In the following schematic diagram for the preparation of hydrogen gas as shown in the figure, what would happen if following changes are made?



(a) In place of zinc granules, same amount of zinc dust is taken in the test tube

(b) Instead of dilute sulfuric acid, dilute hydrochloric acid is taken

(c) In place of zinc, copper turnings are taken

(d) Sodium hydroxide is taken in place of dilute sulfuric acid and the tube is heated.

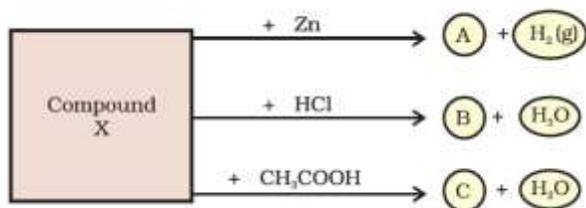
16. For making cake, baking powder is taken. If at home your mother uses baking soda instead of baking powder in cake,

(a) how will it affect the taste of the cake and why?

(b) how can baking soda be converted into baking powder?

(c) what is the role of tartaric acid added to baking soda?

17. A metal carbonate X on reacting with an acid gives a gas which when passed through a solution Y gives the carbonate back. On the other hand, a gas G that is obtained at anode during electrolysis of brine is passed on dry Y, it gives a compound Z, used for disinfecting drinking water. Identity X, Y, G and Z.
18. Identify the compound X on the basis of the reactions given below. Also, write the name and chemical formulae of A, B and C.



19. A dry pellet of a common base B, when kept in open absorbs moisture and turns sticky. The compound is also a by-product of chloralkali process. Identify B. What type of reaction occurs when B is treated with an acidic oxide? Write a balanced chemical equation for one such solution.
20. A sulfate salt of Group 2 element of the Periodic Table is a white, soft substance, which can be molded into different shapes by making its dough. When this compound is left in open for some time, it becomes a solid mass and cannot be used for molding purposes. Identify the sulfate salt and why does it show such a behavior? Give the reaction involved.
21. What are strong and weak acids? In the following list of acids, separate strong acids from weak acids. Hydrochloric acid, citric acid, acetic acid, nitric acid, formic acid, sulfuric acid.

22. Fill in the missing data in the following table

Name of the salt	Formula	Salt obtained from	
		Base	Acid
(i) Ammonium chloride	NH ₄ Cl	NH ₄ OH	—
(ii) Copper sulphate	—	—	H ₂ SO ₄
(iii) Sodium chloride	NaCl	NaOH	—
(iv) Magnesium nitrate	Mg(NO ₃) ₂	—	HNO ₃
(v) Potassium sulphate	K ₂ SO ₄	—	—
(vi) Calcium nitrate	Ca(NO ₃) ₂	Ca(OH) ₂	—

Also indicate whether the above salts are *acidic*, *basic* or *neutral*.

23. How would you distinguish between baking powder and washing soda by heating?
24. In one of the industrial processes used for manufacture of sodium hydroxide, a gas X is formed as by product. The gas X reacts with lime water to give a compound Y which is used as a bleaching agent in chemical industry. Identify X and Y giving the chemical equation of the reactions involved.
25. A student prepared solutions of (i) an acid and (ii) a base in two separate beakers. She forgot to label the solutions and litmus paper is not available in the laboratory. Since both the solutions are colorless, how will she distinguish between the two?
26. What happens when nitric acid is added to egg shell?
27. What will be the action of the following substances on litmus paper?
Dry HCl gas, Moistened NH₃ gas, Lemon juice, Carbonated soft drink, Curd, Soap solution.
28. When zinc metal is treated with a dilute solution of a strong acid, a gas is evolved, which is utilized in the hydrogenation of oil. Name the gas evolved. Write the chemical equation of the reaction involved and also write a test to detect the gas formed.

BIOLOGY

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1. Which of the following statements about the autotrophs is incorrect?
 - (a) They synthesize carbohydrates from carbon dioxide and water in the presence of sunlight and chlorophyll
 - (b) They store carbohydrates in the form of starch
 - (c) They convert carbon dioxide and water into carbohydrates in the absence of sunlight
 - (d) They constitute the first trophic level in food chains
2. In which of the following groups of organisms, food material is broken down outside the body and absorbed?
 - (a) Mushroom, green plants, *Amoeba*
 - (b) Yeast, mushroom, bread mould
 - (c) *Paramecium*, *Amoeba*, *Cuscuta*
 - (d) *Cuscuta*, lice, tapeworm
3. If salivary amylase is lacking in the saliva, which of the following events in the mouth cavity will be affected?
 - (a) Proteins breaking down into amino acids
 - (b) Starch breaking down into sugars
 - (c) Fats breaking down into fatty acids and glycerol
 - (d) Absorption of vitamins
4. The inner lining of stomach is protected by one of the following from hydrochloric acid. Choose the correct one
 - (a) Pepsin
 - (b) Mucus
 - (c) Salivary amylase
 - (d) Bile
5. Which part of alimentary canal receives bile from the liver?
 - (a) Stomach
 - (b) Small intestine
 - (c) Large intestine
 - (d) Oesophagus
6. A few drops of iodine solution were added to rice water. The solution turned blue-black in colour. This indicates that rice water contains
 - (a) complex proteins
 - (b) simple proteins
 - (c) fats
 - (d) starch
7. In which part of the alimentary canal food is finally digested?
 - (a) Stomach
 - (b) Mouth cavity
 - (c) Large intestine
 - (d) Small intestine
8. Choose the function of the pancreatic juice from the following
 - (a) trypsin digests proteins and lipase carbohydrates
 - (b) trypsin digests emulsified fats and lipase proteins
 - (c) trypsin and lipase digest fats

- (d) trypsin digests proteins and lipase emulsified fats
9. Which of the following statement(s) is (are) true about respiration?
- (i) During inhalation, ribs move inward and diaphragm is raised
 - (ii) In the alveoli, exchange of gases takes place i.e., oxygen from alveolar air diffuses into blood and carbon dioxide from blood into alveolar air
 - (iii) Haemoglobin has greater affinity for carbon dioxide than oxygen
 - (iv) Alveoli increase surface area for exchange of gases
- (a) (i) and (iv) (b) (ii) and (iii)
(c) (i) and (iii) (d) (ii) and (iv)
10. Which is the correct sequence of air passage during inhalation?
- (a) Nostrils → larynx → pharynx → trachea → lungs
 - (b) Nasal passage → trachea → pharynx → larynx → alveoli
 - (c) larynx → nostrils → pharynx → lungs
 - (d) Nostrils → pharynx → larynx → trachea → alveoli
11. Which of the following statement (s) is (are) true about heart?
- (i) Left atrium receives oxygenated blood from different parts of body while right atrium receives deoxygenated blood from lungs
 - (ii) Left ventricle pumps oxygenated blood to different body parts while right ventricle pumps deoxygenated blood to lungs
 - (iii) Left atrium transfers oxygenated blood to right ventricle which sends it to different body parts
 - (iv) Right atrium receives deoxygenated blood from different parts of the body while left ventricle pumps oxygenated blood to different parts of the body
- (a) (i) (b) (ii)
(c) (ii) and (iv) (d) (i) and (iii)
12. What prevents backflow of blood inside the heart during contraction?
- (a) Valves in heart
 - (b) Thick muscular walls of ventricles
 - (c) Thin walls of atria
 - (d) All of the above
13. The filtration units of kidneys are called
- (a) ureter
 - (b) urethra
 - (c) neurons
 - (d) nephrons
14. Oxygen liberated during photosynthesis comes from
- (a) water
 - (b) chlorophyll
 - (c) carbon dioxide
 - (d) glucose
15. The blood leaving the tissues becomes richer in
- (a) carbon dioxide
 - (b) water
 - (c) haemoglobin
 - (d) oxygen
16. The internal (cellular) energy reserve in autotrophs is
- (a) glycogen
 - (b) protein
 - (c) starch

(d) fatty acid

17. The opening and closing of the stomatal pore depends upon

(a) oxygen

(b) temperature

(c) water in guard cells

(d) concentration of CO₂ in stomata

18. Choose the forms in which most plants absorb nitrogen

(i) Proteins

(ii) Nitrates and Nitrites

(iii) Urea

(iv) Atmospheric nitrogen

(a) (i) and (ii) (b) (ii) and (iii)

(c) (iii) and (iv) (d) (i) and (iv)

19. Which is the first enzyme to mix with food in the digestive tract?

(a) Pepsin

(b) Cellulase

(c) Amylase

(d) Trypsin

20. Choose the correct path of urine in our body

(a) kidney → ureter → urethra → urinary bladder

(b) kidney → urinary bladder → urethra → ureter

(c) kidney → ureters → urinary bladder → urethra

(d) urinary bladder → kidney → ureter → urethra

21. Name the following:

(6 marks)

(a) The process in plants that links light energy with chemical energy

(b) Organisms that can prepare their own food

(c) The cell organelle where photosynthesis occurs

(d) Cells that surround a stomatal pore

(e) Organisms that cannot prepare their own food

(f) An enzyme secreted from gastric glands in stomach that acts on proteins.

22. Two green plants are kept separately in oxygen free containers, one in the dark and the other in continuous light. Which one will live longer? Give reasons. (2 marks).

23. If a plant is releasing carbon dioxide and taking in oxygen during the day, does it mean that there is no photosynthesis occurring? Justify your answer. (2 marks)

24. Leaves of a healthy potted plant were coated with vaseline. Will this plant remain healthy for long? Give reasons for your answer. (2 marks)

25. Differentiate between an autotroph and a heterotroph. (3 marks)

26. Differentiate between an artery and a vein. (3 marks)

27. Mention the major events during photosynthesis. (3 marks)

28. Why is blood circulation in human heart called double circulation? (5 marks)

29. What is the advantage of having four chambered heart? (5 marks)

30. In each of the following situations what happens to the rate of photosynthesis? (5marks)

(a) Cloudy days

(b) No rainfall in the area

(c) Good manuring in the area

(d) Stomata get blocked due to dust

MATHEMATICS

CLASS: X

TOPIC: REAL NUMBERS

SUBJECT: MATHEMATICS

1. If $7 \times 5 \times 3 \times 2 + 3$ is composite number? Justify your answer
2. Show that any positive odd integer is of the form $4q + 1$ or $4q + 3$ where q is a positive integer
3. Prove that $\sqrt{2} + \sqrt{5}$ is irrational
4. Use Euclid's Division Algorithms to find the H.C.F of
a) 135 and 225 (45)
b) 4052 and 12576 (4)
c) 270, 405 and 315
5. Prove that $5 - 2\sqrt{3}$ is an irrational number
6. Find the HCF and LCM of 26 and 91 and verify that $\text{LCM} \times \text{HCF} = \text{Product of two numbers}$ (13,182)
7. Explain why $\frac{29}{2^3 \times 5^3}$ is a terminating decimal expansion
8. given that $\text{LCM}(77, 99) = 693$, find the HCF (77, 99) (11)
9. Find the greatest number which exactly divides 280 and 1245 leaving remainder 4 and 3 (138)
10. Prove that $\sqrt{2}$ is irrational
11. The LCM of two numbers is 64699, their HCF is 97 and one of the numbers is 2231. Find the other (2813)
12. If $\text{HCF}(6, a) = 2$ and $\text{LCM}(6, a) = 60$ then find a (20)
13. Two numbers are in the ratio 15: 11. If their HCF is 13 and LCM is 2145 then find the numbers (195,143)
14. Express 0.363636..... in the form a/b (4/11)
15. Find the HCF 52 and 117 and express it in form $52x + 117y$
16. Write the HCF of smallest composite number and smallest prime number
17. Write whether $\frac{2\sqrt{45} + 3\sqrt{20}}{2\sqrt{5}}$ on simplification give a rational or an irrational number

SOCIAL SCIENCE

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Make a project report on any issue in your locality/ neighborhood related to environment.

Your report should cover the following-

1. Describe the nature of the environment problem.
2. Collect evidence/data to show the severity of the problem in your locality.
3. Examine and explain possible reasons for this.
4. How does this adversely affect the environment and society?
5. Find five possible solutions / steps to reduce or eliminate the problem.

As per CBSE guidelines –

1. Every student has to compulsorily undertake one project.
2. The overall objective of the project work is to help students gain an insight and pragmatic understanding of the theme and see all the Social Science disciplines from an interdisciplinary perspective. It should also help in enhancing life skills of the students.
3. The project report should be handwritten.

4. If required the student may go out for collecting data and use different secondary resources to prepare the project. If possible various forms of art may be integrated in the project work.
5. The distribution of marks over different aspects related to project work is as follows:

S.No	Aspects	Marks
1.	Content accuracy, originality and analysis	2
2.	Presentation and Creativity	2
3.	Viva Voce	1

- Project Report to be submitted in the first SST ILP after holidays.

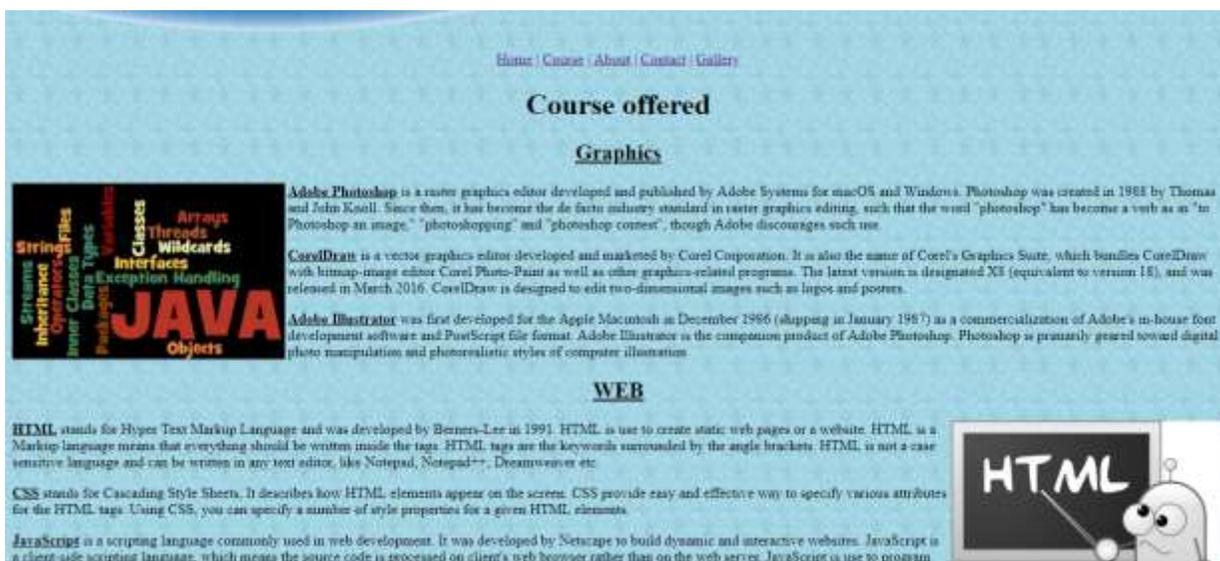
COMPUTER APPLICATIONS

1. Create Simple website with 4 pages (Home, About, Gallery, Course)

Home:



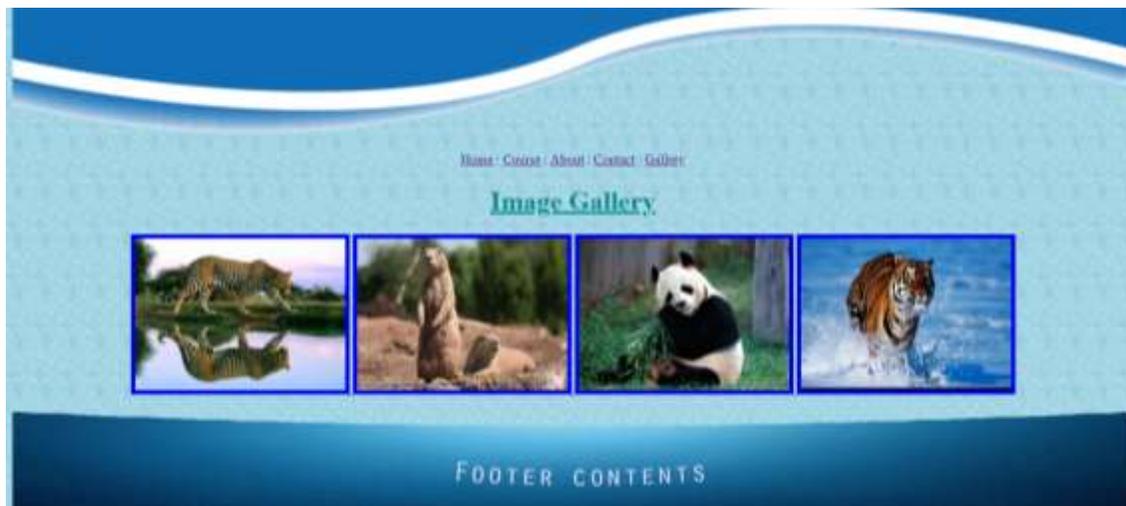
Course Page



About Page



Gallery page



2. Periodic table using table tag:

The Periodic Table																	
1 H																	2 He
3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne
11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
55 Cs	56 Ba	57-71 La	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn
87 Fr	88 Ra	89-103 Ac	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og
72 La	73 Ce	74 Pr	75 Nd	76 Pm	77 Sm	78 Eu	79 Gd	80 Tb	81 Dy	82 Ho	83 Er	84 Tm	85 Yb	86 Lu			
89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr			